CONCEPT: VOLTAMMETRY

Galvanic/Voltaic Cell: A spontaneous cell that _____ or _____ electricity.

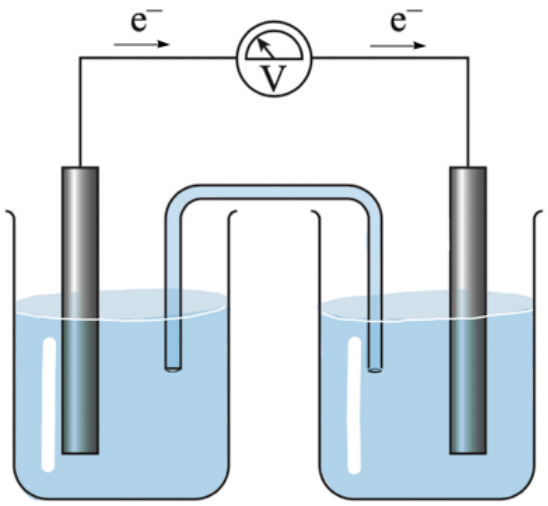
Ionization Energy _____

Anode _____

Producing ⇧ Voltage

[Anode] _____

[Cathode] _____



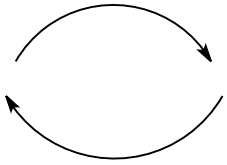
_____ (−)

_____ (+)

Galvanic/Voltaic Cell

Electron Affinity _____

Cathode _____



Cathode : $3\,Cu^{2+}(aq) + \cancel{6\,e^-} \rightleftharpoons 3\,Cu\,(s)$

Anode : $2\,Cr\,(s) \rightleftharpoons 2\,Cr^{3+}(aq) + \cancel{6\,e^-}$

Reduction Half-Reactions	E° (V)
$F_2\,(g) + 2e^- \rightleftharpoons 2F^-$	2.890
$O_3\,(g) + 2H^+ + 2e^- \rightleftharpoons O_2\,(g) + H_2O$	2.075
$MnO_4^- + 8H^+ + 5e^- \rightleftharpoons Mn^{2+} + 4H_2O$	1.507
$Ag^+ + e^- \rightleftharpoons Ag\,(s)$	0.799
$Cu^{2+} + 2e^- \rightleftharpoons Cu\,(s)$	0.339
$2H^+ + 2e^- \rightleftharpoons H_2\,(g)$	0.000
$Cd^{2+} + 2e^- \rightleftharpoons Cd\,(s)$	-0.402
$K^+ + e^- \rightleftharpoons K\,(s)$	-2.936
$Li^+ + e^- \rightleftharpoons Li\,(s)$	-3.040