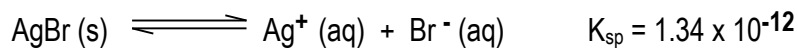
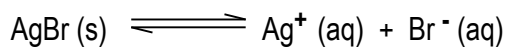


CONCEPT: IONIC STRENGTH

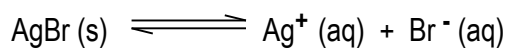
Consider the dissociation of silver bromide, AgBr, in purified water.



Adding 0.10 M NaBr or 0.25 M AgC₂H₃O₂ cause the overall solubility of AgBr to _____ as a result of the _____.



Adding 0.01 M NaClO₄ causes the overall solubility of AgBr to _____ as a result of the **ionic strength**, which is just a measurement of all the ions in the aqueous solution.



Ionic strength represents interactions between the ions in water and the ions of a solution.

$$\mu = \frac{1}{2} \sum c_i z_i^2 = \frac{1}{2} (c_1 z_1^2 + c_2 z_2^2 + \dots)$$

EXAMPLE: Calculate the ionic strength of the following ionic compound.

0.010 M CuSO₃

CONCEPT: IONIC STRENGTH CALCULATIONS

EXAMPLE 1: Calculate the ionic strength for the following ionic compound.



EXAMPLE 2: What is the ionic strength of a solution that is 0.1 M Na_3PO_4 and 0.05 M Na_2HPO_4 ?

PRACTICE: Calculate the ionic strength for the following ionic compound.

