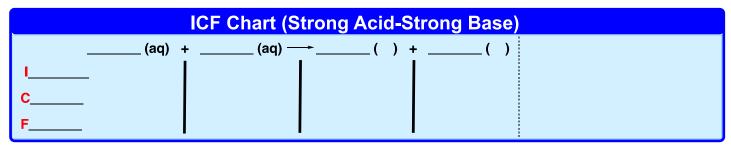
CONCEPT: TITRATIONS: STRONG ACID-STRONG BASE
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This type of titration has either the strong acid or strong base act as the _____ or ____.
When a strong acid and strong base mix together use an ICF (I_____, C____, F____) Chart.

EXAMPLE: Calculate the pH of the solution resulting from the titration of 75.0 mL of 0.100 M HBrO₄ with 55.0 mL of 0.100 M NaOH.

STEP 1: Setup an ICF Chart with both strong species set as _____.



STEP 2: Using the INITIAL ROW, place the given amounts in ______.

□ In this ICF Chart we only care about the ______ & _____.

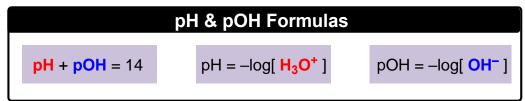
STEP 3: Using the **CHANGE ROW**, looking at the reactants subtract from their initial amounts by the _____ mole amount.

STEP 4: Using the **FINAL ROW**, determine the concentration of the remaining strong species.

□ Divide its final _____ by the total volume used in the chemical reaction.

STEP 5: Recall, [H+] and [OH-] are ______ to the concentration of strong acids and bases, respectively.

 \Box If strong acid remains use its ____ to find pH. \Box If strong base remains use its ____ to find pOH.



CONCEPT: TITRATIONS: STRONG ACID-STRONG BASE
PRACTICE: Calculate the pH of the solution resulting from the mixing of 175.0 mL of 0.250 M HNO ₃ with 75.0 mL of 0.200
M Ba(OH) ₂ .
DDACTICE, Coloulate the nH of the colution regulting from the titration of 110.0 ml of 0.200 M HCl with 220.0 ml of 0.100
PRACTICE: Calculate the pH of the solution resulting from the titration of 110.0 mL of 0.300 M HCl with 330.0 mL of 0.100 M LiOH.