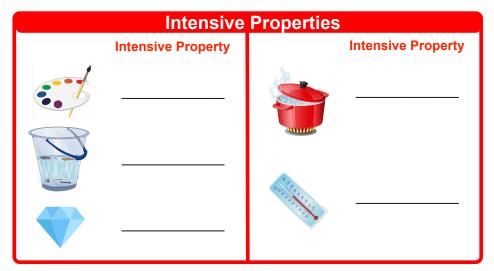
CONCEPT: INTENSIVE VS EXTENSIVE PROPERTIES

- Intensive Properties (______) are those that are _____ of the size or amount of substance present.
 - □ Intensive properties are characterized as _____ properties.



EXAMPLE: Which of the following are examples of intensive properties?

i) Mass

ii) Length

iii) Melting point

iv) Volume

v) Luster

a) i, ii, & iv

b) ii & iii

c) ii, iii & v

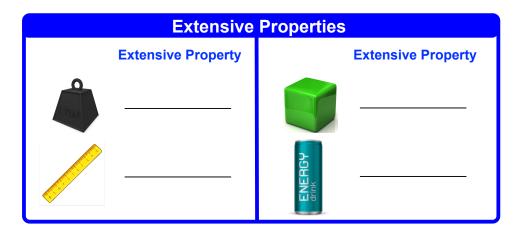
d) iii & v

PRACTICE: Which of the following is <u>not</u> an example of an intensive property?

- a) Elemental Magnetism
- b) Density
- c) Temperature
- d) Milligrams
- e) Freezing Point

CONCEPT: INTENSIVE VS EXTENSIVE PROPERTIES

- Extensive Properties (______) are those that are _____ of the size or amount of substance present.
 - □ Extensive properties are characterized as ______ properties.



EXAMPLE: Which of the following is classified as an extensive property?

- a) Chemical Energy
- b) Electrical Conductivity
- c) Luster
- d) Freezing Point

PRACTICE: Which of the following is **not** an example of an extensive property?

a) Kilograms

b) Gibbs Free Energy

- c) Milliliters
- d) Malleability

PRACTICE: Which of the following is an extensive property of a nitrogen molecule?

- a) Boiling Point
- b) Temperature
- c) Moles
- d) Polarity
- e) Reactivity