






## CONCEPT: INTENSIVE VS EXTENSIVE PROPERTIES

● **Intensive Properties** ( \_\_\_\_\_ ) are those that are \_\_\_\_\_ of the size or amount of substance present.

□ **Intensive properties** are characterized as \_\_\_\_\_ properties.

Intensive Properties	
	<b>Intensive Property</b> _____ _____ _____
	_____ _____ _____
	_____ _____ _____
	<b>Intensive Property</b> _____ _____ _____
	_____ _____ _____

**EXAMPLE:** Which of the following are examples of intensive properties?

i) Mass

ii) Length

iii) Melting point

iv) Volume

v) Luster

a) i, ii, & iv

b) ii & iii

c) ii, iii & v

d) iii & v

**PRACTICE:** Which of the following is not an example of an intensive property?

a) Elemental Magnetism

b) Density





c) Temperature

d) Milligrams

e) Freezing Point

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  - Extensive properties are characterized as \_\_\_\_\_ properties.

Extensive Properties	
 	<b>Extensive Property</b> _____ _____
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**EXAMPLE:** Which of the following is classified as an extensive property?

- a) Chemical Energy      b) Electrical Conductivity      c) Luster      d) Freezing Point

**PRACTICE:** Which of the following is not an example of an extensive property?

- a) Kilograms      b) Gibbs Free Energy      c) Milliliters      d) Malleability

**PRACTICE:** Which of the following is an extensive property of a nitrogen molecule?

- a) Boiling Point  
b) Temperature  
c) Moles  
d) Polarity  
e) Reactivity