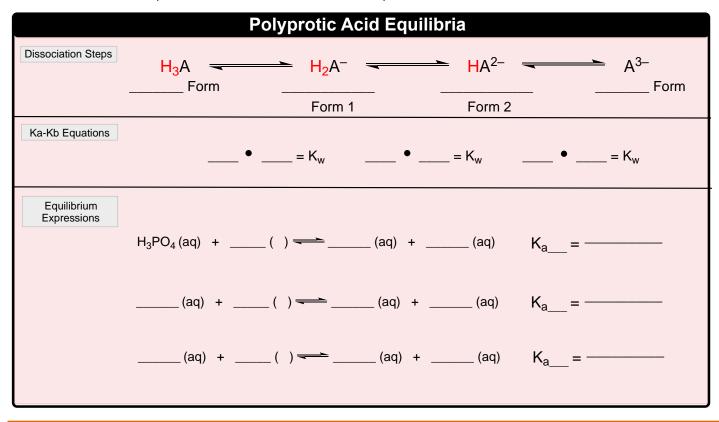
CONCEPT: TRIPROTIC ACIDS AND BASES

Acid Dissociation Constant

- Triprotic Acids (_____) can donate ___ acidic hydrogens and as a result possess three Ka values.
 - □ In terms of Ka magnitude: _____ > ____ > ____
 - K_{a1} deals with donating the acidic proton (H+).
 - Ka2 deals with donating the ____ acidic proton (H+).
 - Ka3 deals with donating the ____ acidic proton (H+).
 - □ The relationships between the Ka values and their respective Kb values are shown as:



EXAMPLE: Provide the dissociation equation associated with the K_{a2} value for the triprotic acid of pyrophosphoric acid, $H_4P_2O_7$.

PRACTICE: Determine the equilibrium expression for the Ka3 value of citric acid, H3C6H5O7?