CONCEPT: INTENSIVE VS EXTENSIVE PROPERTIES

- **Intensive Properties** (_____) are those that are _____ of the size or amount of substance present.
  - Intensive properties are characterized as _____ properties.

Intensive Properties

Intensive Property		Intensive Property	
	_____		_____
	_____		_____
	_____		

EXAMPLE: Which of the following are examples of intensive properties?

i) Mass ii) Length iii) Melting point iv) Volume v) Luster

a) i, ii, & iv b) ii & iii c) ii, iii & v d) iii & v

PRACTICE: Which of the following is **not** an example of an intensive property?

a) Elemental Magnetism

b) Density

c) Temperature

d) Milligrams

e) Freezing Point

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Extensive Properties	
Extensive Property	Extensive Property
 _____	 _____
 _____	 _____

EXAMPLE: Which of the following is classified as an extensive property?

a) Chemical Energy b) Electrical Conductivity c) Luster d) Freezing Point

PRACTICE: Which of the following is **not** an example of an extensive property?

a) Kilograms b) Gibbs Free Energy c) Milliliters d) Malleability

PRACTICE: Which of the following is an extensive property of a nitrogen molecule?

a) Boiling Point

b) Temperature

c) Moles

d) Polarity

e) Reactivity