

CONCEPT: IONIC SALTS (SIMPLIFIED)

When an acid neutralizes a base an ionic compound called a _____ is formed.

- These solutions can be neutral, acidic or basic, based on acid-base properties of the cations and anions formed.

RULES FOR IDENTIFYING YOUR IONS

CATIONS (POSITIVE IONS)

- 1) **Transition Metals**: If your transition metal has a charge of +2 or higher it is acidic. If the charge is less than +2 then it is neutral.

EX:

- 2) **Main-Group Metals**: If your main-group metal has a charge of +3 or higher it is acidic. If the charge is less than +3 then it is neutral.

EX:

- 3) **Positive Amines** are acidic.

EX:

ANIONS (NEGATIVE IONS)

- 1) **NEGATIVE ION**: If you have a negative ion then add an H^+ to it. If you create a weak acid then your negative ion is basic.

EX:

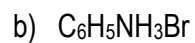
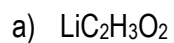
EXAMPLE: Determine if each of the following compounds will create an acidic, basic or neutral solution.

a) NaOCl

b) $PbCl_4$

CONCEPT: IONIC SALTS (SIMPLIFIED)

PRACTICE: Determine if each of the following compounds will create an acidic, basic or neutral solution.



PRACTICE: Determine if each of the following compounds will create an acidic, basic or neutral solution.

