CONCEPT: IONIC SALTS (SIMPLIFIED)

When an acid neutralizes a base an ionic compound called a _____ is formed.

- These solutions can be neutral, acidic or basic, based on acid-base properties of the cations and anions formed.

RULES FOR IDENTIFYING YOUR IONS

CATIONS (POSITIVE IONS)

1) **Transition Metals:** If your transition metal has a charge of +2 or higher it is acidic. If the charge is less than +2 then it is neutral.

   EX:

2) **Main-Group Metals:** If your main-group metal has a charge of +3 or higher it is acidic. If the charge is less than +3 then it is neutral.

   EX:

3) **Positive Amines** are acidic.

   EX:

ANIONS (NEGATIVE IONS)

1) **NEGATIVE ION**: If you have a negative ion then add an H^+ to it. If you create a weak acid then your negative ion is basic.

   EX:

EXAMPLE: Determine if each of the following compounds will create an acidic, basic or neutral solution.

a) NaOCl b) $PbCl_4$

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PRACTICE: Determine if each of the following compounds will create an acidic, basic or neutral solution.

a) $LiC_2H_3O_2$

b) $C_6H_5NH_3Br$

PRACTICE: Determine if each of the following compounds will create an acidic, basic or neutral solution.

a) $Co(HSO_4)_2$

b) $Sr(HSO_3)_2$