

CONCEPT: BRONSTED LOWRY ACIDS & BASES

In 1923, Johannes Brønsted and Thomas Lowry developed a new set of definitions for acids and bases.

According to the Bronsted-Lowry definition, acids are considered _____ and bases are considered _____.

- Unlike Arrhenius acids and bases, they are not limited to aqueous solutions.
- Every Arrhenius acid is a Brønsted-Lowry acid (and likewise for the bases).
- Brønsted-Lowry acids and bases always occur in pairs called _____.

EXAMPLE: Write the formula of the **conjugate base** for the following compound:



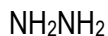
EXAMPLE: Write the formula of the **conjugate acid** for the following compound:



PRACTICE: Write the formula of the **conjugate base** for the following compound:

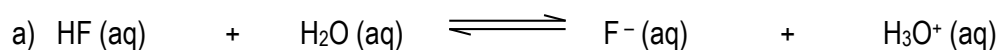


PRACTICE: Write the formula of the conjugate for the following compound:

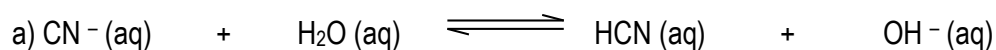


PRACTICE: BRONSTED LOWRY ACIDS & BASES

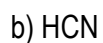
EXAMPLE: Identify the acid, base, conjugate acid and conjugate base in the following reactions:



EXAMPLE: Identify the acid, base, conjugate acid and conjugate base in the following reactions:



PRACTICE: Which of the following is a Bronsted-Lowry acid?



PRACTICE: Determine the chemical equation that would result when carbonate, CO_3^{2-} , reacts with water.

