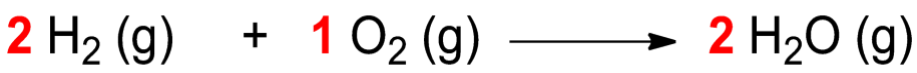
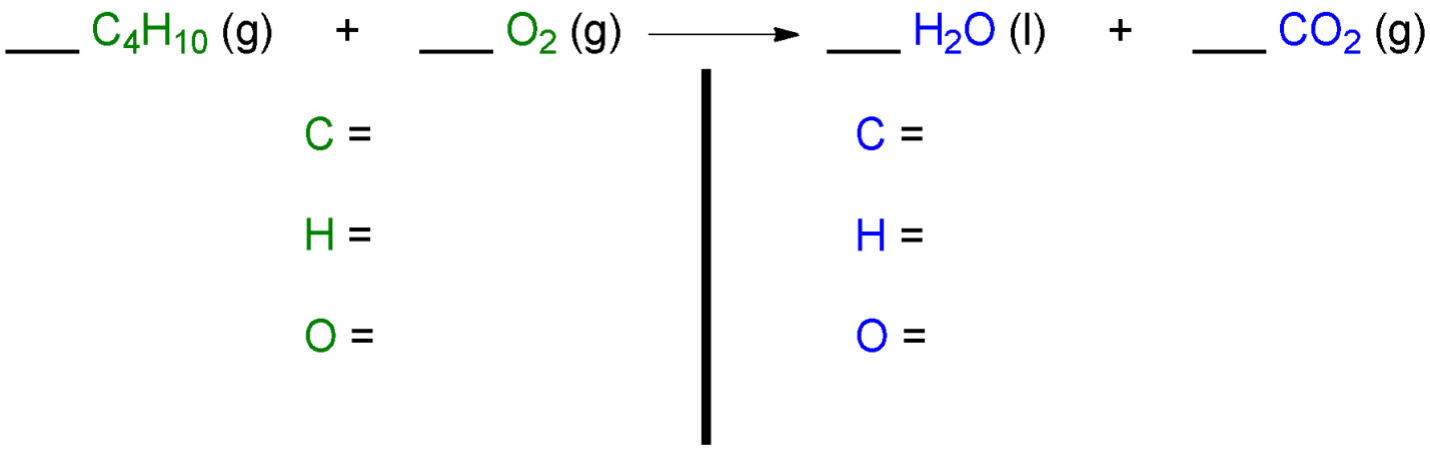
**CONCEPT:** BALANCING CHEMICAL EQUATIONS (SIMPLIFIED)

- When balancing always make sure the _____ and _____ of atoms on both sides of the arrow are equal.
  - □ In a balanced equation the **numbers** are referred to as _____.




**EXAMPLE**: Write the balanced equation for the following by inserting the correct coefficients in the blanks:

**STEP 1:** Set up a list for the elements that are **Reactants** and another list for the elements that are **Products**.




**STEP 2:** Start from the top and going down both lists determine how many of each element is present.

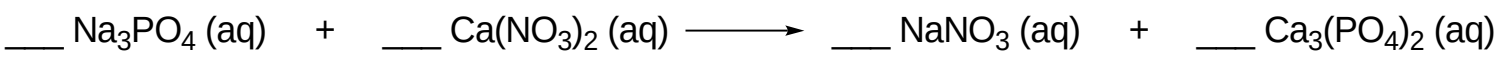
- □ If a polyatomic ion is present on both sides, treat it as a _____ unit.

**STEP 3:** Start from the top and going down both lists begin balancing each element to ensure they match.

- □ Sometimes you may have a decimal or a fraction as a coefficient and so must multiply the equation by _____.

**CONCEPT:** BALANCING CHEMICAL EQUATIONS (SIMPLIFIED)

**PRACTICE:** Write the balanced equation for the following by inserting the correct coefficients in the blanks.




**PRACTICE:** Determine the total sum of the coefficients after balancing the following equation.