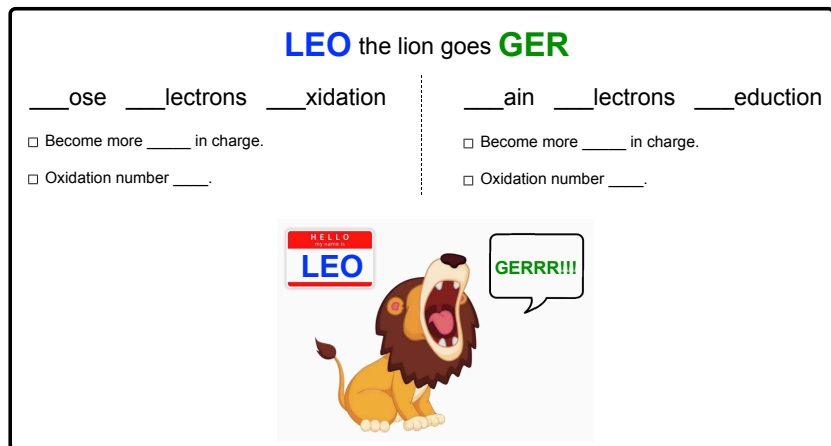


## CONCEPT: REDOX REACTIONS

- **Redox Reactions** (oxidation-reduction reactions) involve transference of an electron(s) between reactants.

### MEMORY TOOL



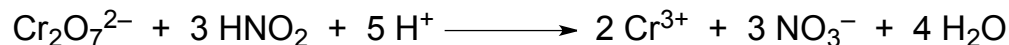
- **Reduction and Oxidation:** the \_\_\_\_\_ of oxidizing and reducing agents.
  - ☐ **Oxidizing Agent:** The element or compound that is \_\_\_\_\_.
  - ☐ **Reducing Agent:** The element or compound that is \_\_\_\_\_.

**EXAMPLE:** Consider the following reaction below when solid lithium reacts with the zinc ion:



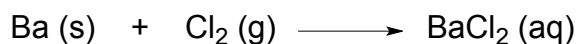
Which reactant is undergoing **oxidation** and which reactant is undergoing **reduction**?

**PRACTICE:** Which element is being reduced in the following reaction?

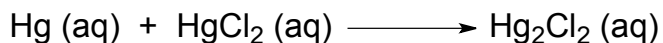


**CONCEPT: REDOX REACTIONS**

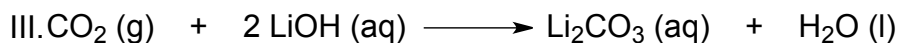
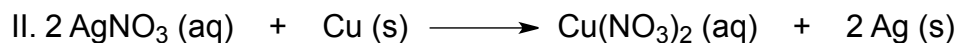
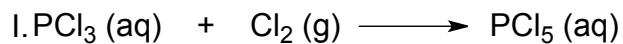
**PRACTICE:** Identify the oxidizing agent and reducing agent from the following redox reaction.



**PRACTICE:** Which element is oxidized and which is reduced in the following reaction?



**PRACTICE:** Which of the following represents an oxidation-reduction reaction?



a) I, II, III, and IV

b) I, II, and III

c) III

d) I and II

e) IV