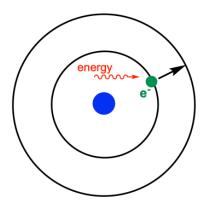
## **CONCEPT: GAMMA EMISSION**

Gamma Radiatio	<b>n</b> is related to the electromagnetic	spectrum. Gamma rays have the h	ighest energy and therefore they
have	wavelength and	frequency.	

A gamma particle can be represented by \_\_\_\_\_\_.

- It causes no change in the atomic mass or atomic number and usually happens with alpha or beta decay.
- Gamma particles have the \_\_\_\_\_\_ ionizing power.
- Gamma particles have the \_\_\_\_\_ penetrating power so thick layers of lead shielding are needed.



**EXAMPLE**: Which of the following represents an element that has experienced a gamma emission?

a. Cl:  $1s^22s^22p^63s^23p^5$ 

b. Be: 1s<sup>2</sup>2s<sup>2</sup>

c. Na: 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3p<sup>1</sup>

d. N:  $1s^22s^22p^3$