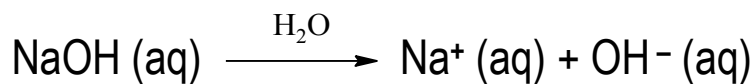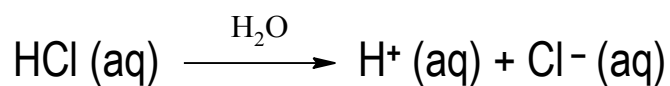**CONCEPT:** pH OF STRONG ACIDS & BASES

STRONG ACIDS & BASES are considered _____ Electrolytes so they ionize completely in water.

$$HCl\ (aq)\ \xrightarrow{\ H_2O\ }\ H^+\ (aq) + Cl^-\ (aq)$$

$$NaOH\ (aq)\ \xrightarrow{\ H_2O\ }\ Na^+\ (aq) + OH^-\ (aq)$$

EXAMPLE: Calculate the pH of a 0.0782 M solution of CaH_2.

EXAMPLE: Calculate the pH of a 0.000550 M HBr solution to the correct number of significant figures.

a) 3.3

b) 3.26

c) 3.260

d) 3.2596

e) All are correct

PRACTICE: Calculate the pH of 50.00 mL of 4.3 x 10^{-7} M H_2SO_4.