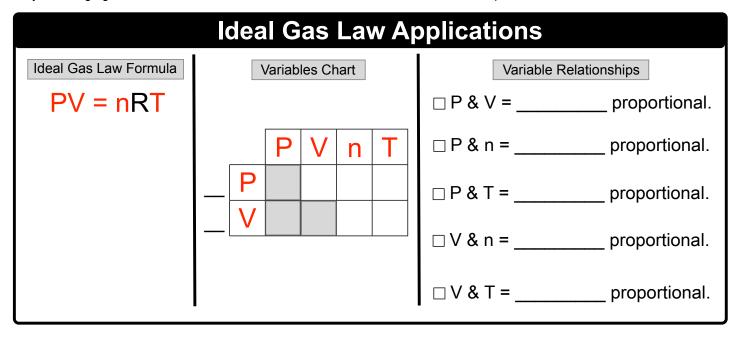
CONCEPT: THE IDEAL GAS LAW APPLICATIONS

• By rearranging *The Ideal Gas Law* we can establish direct and inverse relationships between its variables.



EXAMPLE: If the number of moles (n) inside a container were tripled while keeping the pressure (P) constant, what will happen to the volume (V)?

- a) It will remain the same
- b) It will be cut by a third
- c) It will triple
- d) It will double

PRACTICE: The relationship between the partial pressure of a gas (P) and the number of moles of that gas (n) is best represented by which of the following graphs?

