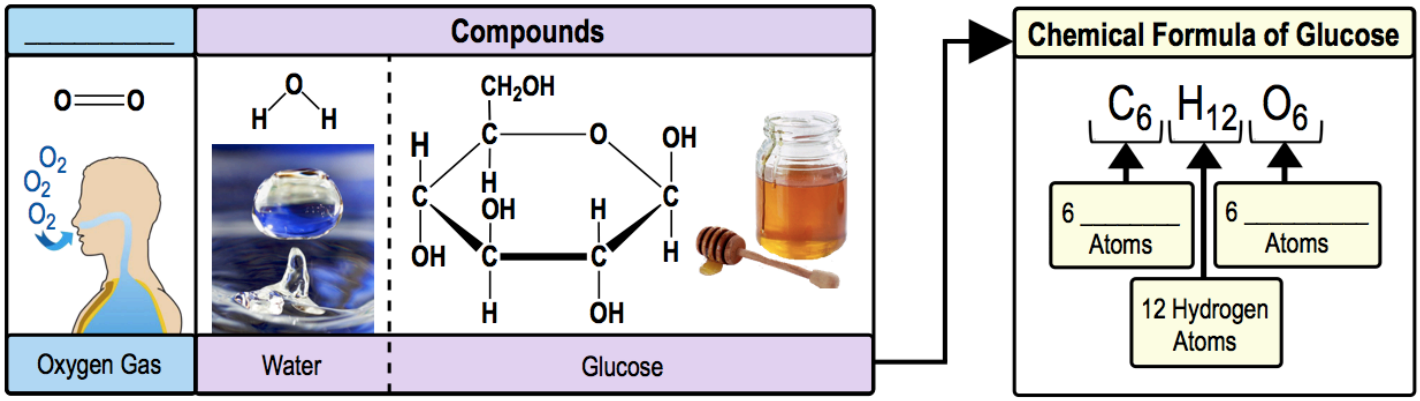
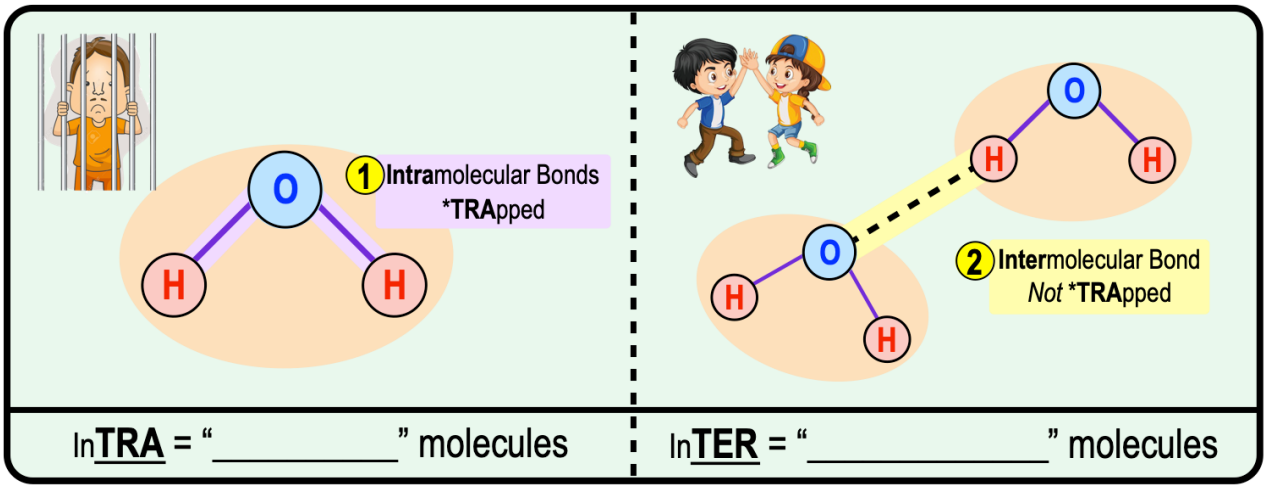
CONCEPT: INTRODUCTION TO CHEMICAL BONDING

- *Chemical* _____: attractive forces *between* atoms, holding them *together* to form *molecules* & *compounds*.
  - □ _____: substance containing ≥ 2 *chemically* bound _____ (ex. O_2).
    - □ _____: *comp*licated molecule composed of ≥ 2 different _____ (ex. H_2O).
  - □ *Chemical Formula*: reveals the _____ & _____ of atoms in a molecule (ex. $C_6H_{12}O_6$).



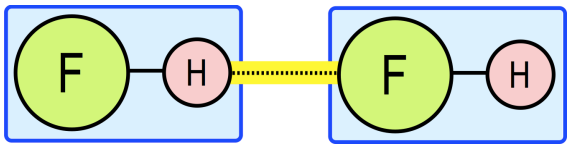
Intramolecular vs. Intermolecular Bonds

- Bonds between atoms can either be _____-molecular or _____-molecular:
  1. *Intramolecular Bonds*: interactions between atoms within the _____ *molecule*.
  2. *Intermolecular Bonds*: interactions between atoms of _____ molecules.



EXAMPLE: The Hydrogen Fluoride (HF) molecules below are interacting with each other through which types of bonds?

a) Intramolecular Bonds.
b) Extramolecular Bonds.
c) Intermolecular Bonds.
d) None of the Above.



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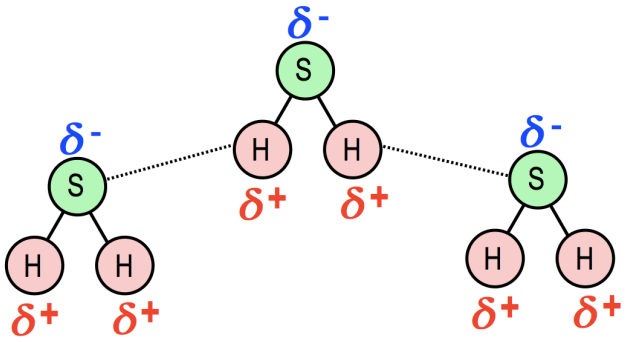
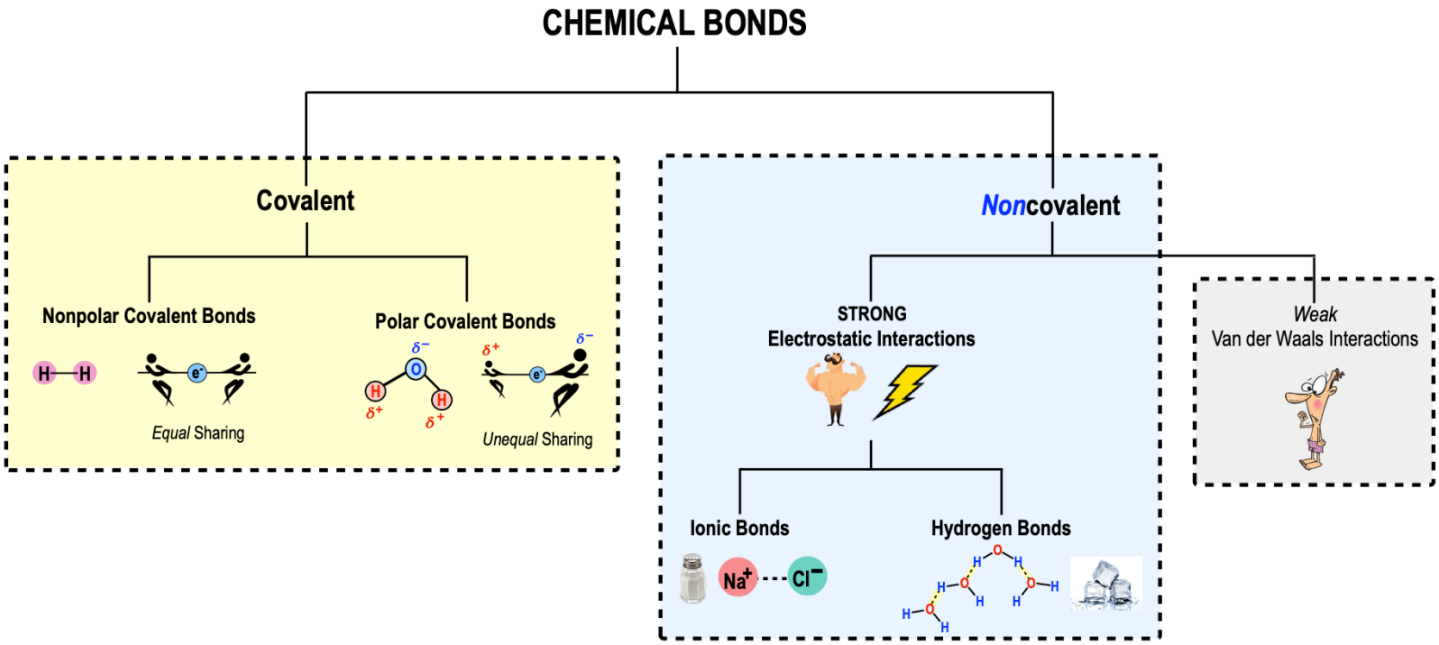
PRACTICE: Appropriately label all of the chemical bonds in this image as either intramolecular or intermolecular.



Map of the Lesson on Chemical Bonding



PRACTICE: According to the map above, which of the following are types of covalent bonds?

a) Polar.
b) Van der Waals.
c) Ionic.
d) Hydrogen.
e) None of the above.