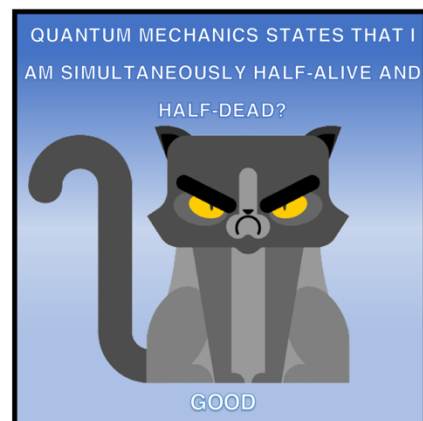
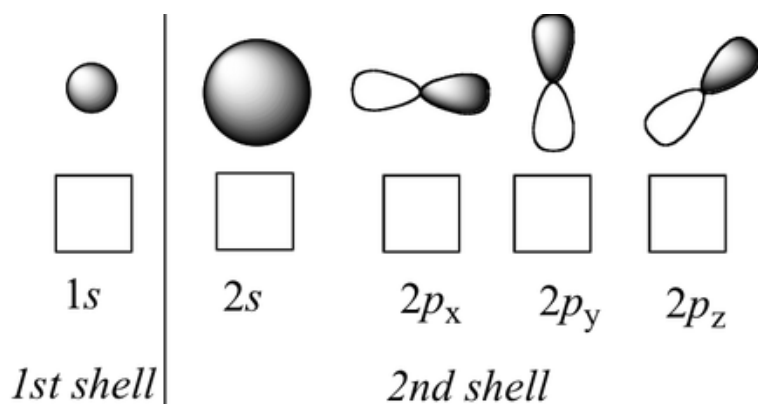


CONCEPT: WAVE FUNCTION

Quantum Mechanics states that electrons behave both as particles and as _____.

- The Heisenberg Uncertainty Principle states that we cannot simultaneously know an electron's speed and _____
 - Equations called wave functions correspond to the energy state of a given electron _____
 - The *relative probability* of finding an electron can be derived from the wave function _____
 - The 3-D plot of the _____ is called an atomic _____: where the chance of finding electrons is **high**.



As with any type of wave, wave functions have the ability to _____ with each other upon meeting.

- This can occur either **constructively** or **destructively**

EXAMPLE: H₂ Molecular Orbitals

