CONCEPT: DISSOCIATION CONSTANT AND pKa

●In general chemistry, we used pH to measure _____

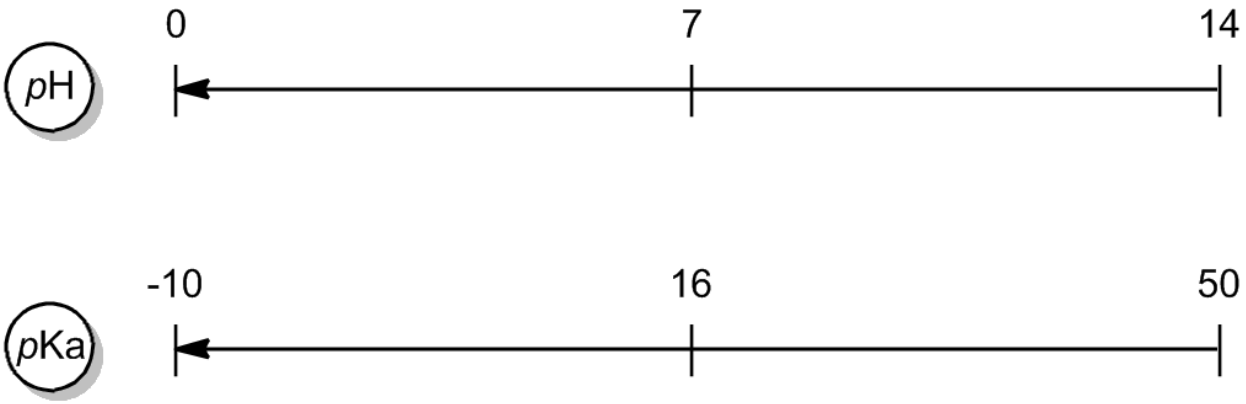
● In organic chemistry, we use _____ to measure the tendency for a molecule to _____

□ Strong acids have a _____ dissociation constant (they _____ dissociate in aqueous solution)

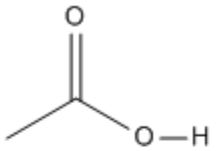
□ Weak acids have a _____ dissociation constant (they _____ dissociate in aqueous solution)

● *p* = $K_a = \left(\frac{[H+][A-]}{[HA]}\right)$ (dissociation constant)

Therefore, the ____ Ka, the _____ the pKa

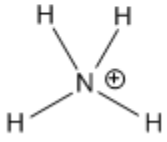


EXAMPLE: Calculate the pK_a's of the following acids and indicate which is the stronger acid.



Acetic Acid

Dissociation Constant = **1.75×10^{-5}**



Ammonium

Dissociation Constant = **5.8×10^{-10}**