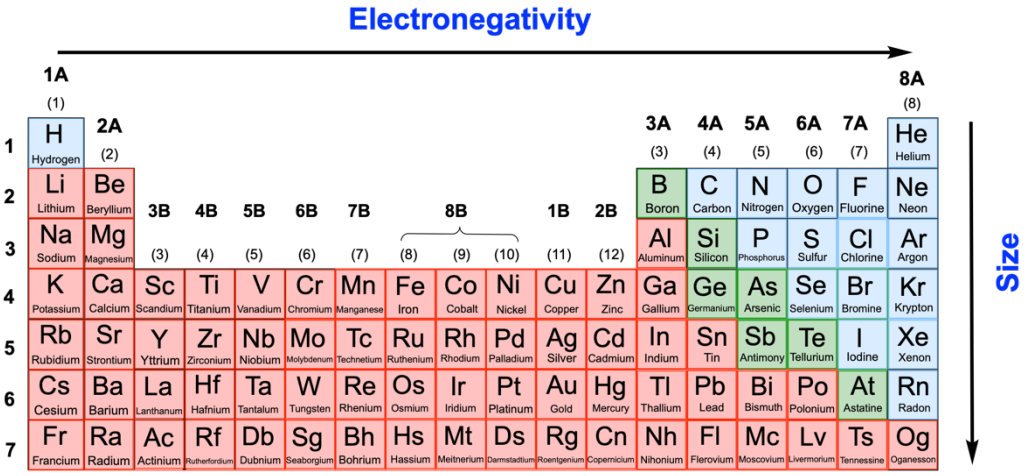
CONCEPT: LEAVING GROUPS

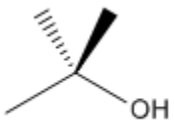
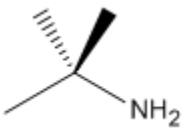
Leaving groups *break a bond* with the electrophile to make it reactive. They are molecules that will remain stable after accepting an extra electron pair.

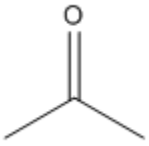
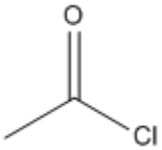
□ We use ***factors affecting acidity*** to determine which atoms will be most stable after gaining extra electrons.

- Recall that the element effect consists of two trends:

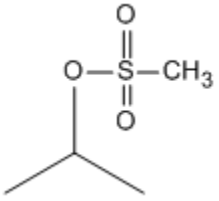
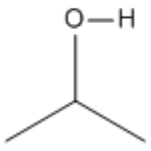


EXAMPLE: Predict which of the following pairs of electrophiles possesses the best leaving group

a.  and 

b.  and 

c. $CH_3—F$ and $CH_3—I$

d.  and 

□ Due to their high electronegativity, _____ _____ will be the primary leaving groups for this chapter.